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Specific Heat Capacity

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300 at Harvard University. Name:
_ Date: _ Student Exploration:
Calorimetry Lab Vocabulary:
calorie, calorimeter, joule, specific

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B. Suppose a snack bar is burned in a calorimeter and heats 2,000 g water by 20 °C. How much heat energy was released? (Hint: Use

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the specific heat equation.) Give your answer in both joules and calories.

~~Student Exploration Calorimetry
Lab (ANSWER KEY)~~

Student Exploration: Calorimetry
Lab Vocabulary: calorie,

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calorimeter, joule, specific heat capacity Prior Knowledge

Questions (Do these BEFORE

using the Gizmo.) 1. The Latin

word calor means "heat," and

meter comes from the Greek

word meaning "to measure."

What do you think a calorimeter

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does? Measures heat 2. Where
have you heard the word calorie
before? ? What do you think a
calorie

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Exploration Calorimetry Lab ...
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Everett Community College
Tutoring Center Student Support
Services Program $C_p(\text{H}_2\text{O}) =$
 $4.184 \text{ J / g } \cdot \text{ }^\circ\text{C}$ $H = mC_p \Delta T$ 1) A
compound is burned in a bomb
calorimeter that contains 3.00 L
of water. If the combustion of
0.285 moles of this compound

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causes the temperature of the
water to

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Substitution into equation 3 yields equation 4: $(4) q_{\text{water}} = m_{\text{water}} \cdot s_{\text{water}} \cdot \Delta T_{\text{water}}$.

Finally, equations 1, 2 and 3 can be combined into the calorimetry

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Equation: $q_{\text{sys}} = -m_{\text{water}} \cdot s_{\text{water}} \cdot \Delta T_{\text{water}} = -m_{\text{water}} \cdot (4.18 \text{ J/g} \cdot ^\circ\text{C}) \cdot \Delta T_{\text{water}}$.

~~Lab 4 — Calorimetry~~

Correct Answer: C. The final temperature of the lead-water system will be lower than the final

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temperature of the copper-water system. A blacksmith heats a 1,540 g iron horseshoe to a temperature of 1445°C before dropping it into 4,280 g of water at 23.1°C .

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Soda Can Calorimeter Energy
Content of Food ... This lab should
be performed in a well-ventilated

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room. Wash hands thoroughly ...
(kilocalories) by dividing the
answer above by 1000. $265 \text{ cal.} \div$
 $1000 = 0.265 \text{ Cal.}$ 4. Determine
how much of the food burned by
subtracting the final mass of the
cork/pin/food assembly from the
initial mass.

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~~Soda Can Calorimeter Flinn
Scientific~~

6.03 Calorimetry Lab Report By;
Selina Pfuner CALCULATIONS P2

Unknown Metals $q[\text{water}] = m \times c$
 $\times \Delta T$ $m = 24.5 \text{ g}$ $c = 4.18$
 $\Delta T = 29.1 - 25.2 = 3.9$

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~~Calorimetry Lab~~
 $\Delta T = \text{final temp} - \text{initial temp}$
 $24.5 \times 4.18 \times 3.9 = 399\text{J}$ multiply
all together equal $q[\text{water}] A B C$
 $-399\text{J} = 25.605[29.1 - 100.5]$

~~6.03 Calorimetry Lab Report by
Selina Pfuner Prezi~~
Calorimetry Lab Report – Science

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lab reports are developed to communicate the findings of research, in a method that is clear to readers. You need to not forget to include any additional details, which may be useful for readers. Composing a great science lab sample is essential if

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Open the Calorimetry Lab Activity
(Section 6.6, Figure 6.11) to
complete this worksheet. 1.

Determine the masses of the cup,
the water, and the gold block.

Object	Cup	Water	Gold	Mass (g)
	4.0g	60.0g	38.69	2.

Measure the
initial temperature of the water.

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Solved: Chapter 6 Calorimetry
Worksheet Key Open

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Question: Name: LAB #4 -
CALORIMETRY Calorimetry Is The
Practice Of Measuring The Heat
Released By A Chemical Reaction
By Observing The Effect Of That
Heat On A Substance Of Known

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Specific Heat (usually Water). If A
Known Mass Of Water Undergoes
A Measurable Temperature
Change As A Result Of A Chemical
Reaction, The Amount Of Heat
Released By The Reaction Can ...

~~Solved: Name: LAB #4~~

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Prac ...

Question: DATA SHEET For
CALORIMETRY EXPERIMENTS A)
SPECIFIC HEAT Of A METAL METAL
UNKNOWN Number Trial 2 (50 MI
Water) 13.4502 23.9235 100.0
5.0250 55.020 Trial 1 (40 MI

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Water) Mass Of Empty Test Tube
(tare) 13.4502 Mass Of Tube +
Metal 23.7502 Temperature Of
Boiling Water 100.0 Mass Of
Empty Calorimeter 5.0250 Mass
Of Calorimeter + Water 45.0250
Initial Temperature ...

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temperature values to see effects on the system. One or any combination of the substances can be mixed with water. A dynamic graph (temperature vs. time) shows temperatures of the individual substances after mixing.

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December 06, 2012 Key Equaon
(try to answer number 13 and if
you are not sure, ask!) 13.To
change the proporonality sign, α ,

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to an equality (= sign), we need
to introduce a proportionality
constant....

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metallic copper from copper ore
you make yourself Analyze the
makeup of seawater, bone, and
other common substances
Synthesize oil of wintergreen from

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